Unit 3 – Bonding

IPOD Questions

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IPOD # 17

Name or Write a Formula for the following examples…

- $Zn(OH)_2$
- Strontium Sulfite
- Tin (II) Sulfide
- Aluminum Carbonate
- CoP
- Magnesium Bromate
- Fe_2O_3
- $Ga(NO_3)_3$
- Silver (I) Iodide
- MnO

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IPOD # 18

Show how an ionic bond forms by completing the following:

Metal	Valence shell	e^- dot symbol	Ion formed	Valence shell of ion	Chemical name and formula of compound
aluminum					
Nonmetal	**Valence shell**	**e^- dot symbol**	**Ion formed**	**Valence shell of ion**	
bromine					
Bonding model					

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IPOD # 19

Name or Write a Formula for the following examples…

- K_3PO_4
- Sodium nitride
- Boron trichloride
- HI
- N_2O_5
- Calcium sulfate
- Acetic acid
- $HBrO_2$
- Lead (II) hydroxide
- Sulfuric acid
- KBr
- CCl_4
- Cu_2SO_4
- Dinitrogen tetrahydride

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IPOD # 20

Draw a Lewis Dot Structure for the Following…

- CH_2Cl_2
- Nitrogen trichloride
- Boron triiodide
- Dihydrogen monosulfide
- Lead (II) hydroxide
- Carbon dioxide

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IPOD # 21

Complete the Following Table…

Name	Formula	Lewis Dot Structure	Molecular Shape VSEPR Shape	Bond Polarity Use EN differences to calculate	Molecule Type polar or nonpolar based on molecule symmetry	Intermolecular Attractions London dispersion, dipole, hydrogen bonding?
silicon tetrachloride						
Dihydrogen monoselenide						
	NH_3