

Chemistry
Chapter 6 – Circle the Trend

Name _____
Date _____

Circle the element with:

1. The HIGHER ionization energy

A. Mg or S

B. Mg or Sr

C. Br or K

D. Br or I

E. Li or Rb

F. As or N

G. O or Te

H. Cu or Fe

2. The HIGHER electronegativity

A. Sr or Mg

B. Cl or Br

C. K or Ge

D. P or Sb

E. Ag or Sr

F. Se or O

G. Ga or B

H. Cs or Na

3. The LARGER radius

A. Li or K

B. I or F

C. Ca or Ga

D. O or S

E. Co or Ir

F. Ne or Xe

G. Cs or Na

H. Sn or C

4. The LOWER ionization energy

A. N or As

B. Br or F

C. Ca or Be

D. Al or Cl

E. Cs or Po

F. N or P

5. The LOWER electronegativity

A. Rb or Sn

B. Br or At

C. Ca or Ni

D. Mg or S

E. Cl or P

F. Al or Ga

6. The SMALLER radius

A. Cs or K

B. Ba or Ca

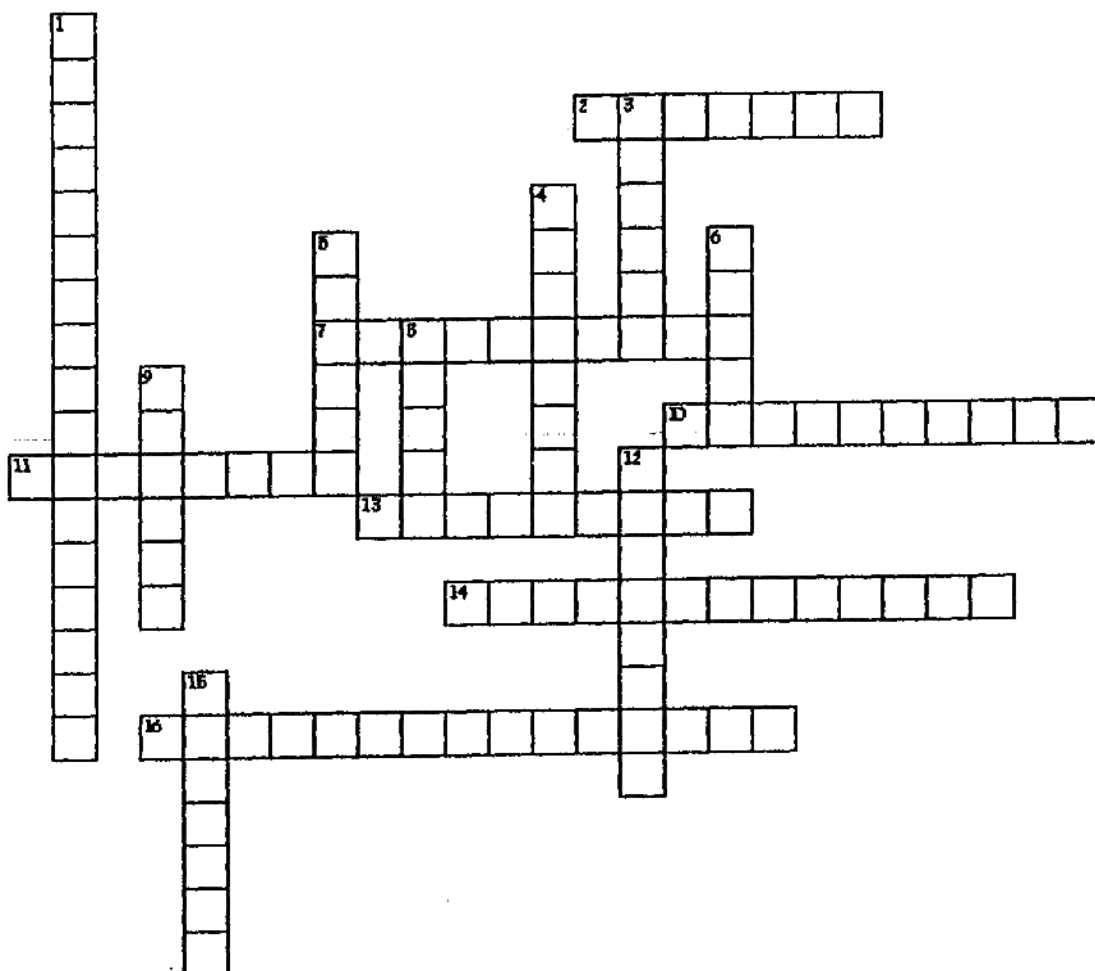
C. N or As

D. Cl or I

E. Cr or Cu

F. Ag or Mo

Chemical Periodicity



Across

2. positively charged ions that have lost one or more electrons
7. energy required to remove an electron from an atom
10. metals with their s & d sublevels partially full
11. nonmetallic elements of group 7A; including F, Cl, Br, I & At
13. Russian chemist who constructed the 1st periodic table
14. half the distance between nuclei of two like atoms
16. elements with s or p sublevels partially full

Down

1. tendency of an element to attract electrons
3. metals in group 1A; including Li, Na, K, Rb, Cs, Fr
4. metals in group 2A; including Be, Mg, Ca, Sr, Ba, Ra
5. negatively charged ions that have gained one or more electrons
6. transition metals with s & f sublevels partially full
8. gases have filled s and p sublevels; the most stable elements
9. vertical columns on the periodic table
12. patterns in the chemical & physical properties of elements
15. horizontal rows across the periodic table