

Procedures

1. Look at the reactions below.
2. Identify the reaction type as **Double Replacement, Single Replacement, Decomposition, Synthesis, Combustion (or a combination of 2 reactions)**.
3. Write formulas for the reactants in each scenario, predict the products that formed if a reaction was observed and balance the overall equation.

Reaction 1 Type: _____

zinc + hydrochloric acid \rightarrow

(if a reaction occurs, assume zinc (II) for the purpose of product formula writing)

Reaction 2 Types: _____ **then** _____

a) copper (II) sulfate + sodium hydroxide \rightarrow

b) copper (II) hydroxide \rightarrow copper (II) oxide + water

Reaction 3 Types: _____ **then** _____

a) potassium chlorate \rightarrow potassium chloride + oxygen gas

b) $C_{12}H_{22}O_{11}$ (sucrose) + oxygen gas \rightarrow

Reaction 4 Type: *both reactions are* _____

a) manganese dioxide →

b) potassium chlorate → potassium chloride + oxygen gas

Reaction 5 Type: _____

Lead (II) nitrate + potassium iodide →

Reaction 6 Type: _____

C₂H₅OH (ethanol) + oxygen gas →