

Procedures

1. Identify the reaction type as either **Double Replacement** or **Single Replacement**.
2. Write formulas for the reactants in each scenario.
3. Use the reference sheet (activity series of metals & solubility rules) to determine if a reaction will occur. If one does, identify the products that formed.
4. Balance the overall equation.

Reaction 1

Reaction Type: _____

potassium chloride + silver (I) nitrate →

Reaction 2

Reaction Type: _____

magnesium + hydrochloric acid →

Reaction 3

Reaction Type: _____

zinc + calcium nitrate →

(if a reaction occurs, assume zinc (II) for the purpose of product formula writing)

Reaction 4

Reaction Type: _____

sodium phosphate + silver (I) nitrate →

Reaction 5

Reaction Type: _____

sodium chloride + calcium nitrate →

Reaction 6

Reaction Type: _____

iron + hydrochloric acid →

(if a reaction occurs, assume iron (III) for the purpose of product formula writing)