

Procedures

1. In a test tube, combine a few drops of each reactant. In the case where solid metals are needed, place the solid in the test tube and then a few drops of the other reactant. Record the observations for each reaction.
2. Identify the reaction type as either **Double Replacement** or **Single Replacement**.
3. Write formulas for the reactants in each scenario, predict the products that formed if a reaction was observed and balance the overall equation.

Reaction 1

Observations:

Reaction Type: _____

potassium chloride + silver (I) nitrate →

Reaction 2

Observations:

Reaction Type: _____

magnesium + hydrochloric acid →

Reaction 3

Observations:

Reaction Type: _____

zinc + calcium nitrate →

(if a reaction occurs, assume zinc (II) for the purpose of product formula writing)

Reaction 4**Observations:****Reaction Type:** _____

sodium phosphate + silver (I) nitrate →

Reaction 5**Observations:****Reaction Type:** _____

sodium chloride + calcium nitrate →

Reaction 6**Observations:****Reaction Type:** _____

iron + hydrochloric acid →

(if a reaction occurs, assume iron (III) for the purpose of product formula writing)