

Procedures

1. Predict the products of the following double replacement reactions. Using the solubility rules as a reference, determine if this reaction will occur (*does a solid, gas or water form?*).
2. If the reaction occurs, balance the equation.
3. On the watch glass, combine a few drops of each reactant. ***The observations for each reaction have been recorded for you to reference.***

1) sodium phosphate + magnesium nitrate

Observations:

2 clear liquids form a cloudy white substance when reacted

2) copper (II) sulfate + barium nitrate

Observations:

2 clear liquids form a cloudy blue substance when reacted

3) potassium chloride + silver (I) nitrate

Observations:

2 clear liquids form a cloudy white substance when reacted

4) sodium chloride + calcium nitrate

Observations:

2 clear liquids react and no observable solid product is formed

5) sodium phosphate + silver nitrate

Observations:

2 clear liquids form a cloudy yellow substance when reacted

6) potassium chloride + sodium hydroxide

Observations:

2 clear liquids react and no observable solid product is formed

7) copper (II) sulfate + calcium nitrate

Observations:

2 clear liquids react to form a cloudy product