Academic Chemistry **Name _____**
Double Replacement Reactions Lab **Date_____ Block____**

Procedures

1. Predict the products of the following double replacement reactions. Using the solubility rules as a reference, determine if this reaction will occur (*does a solid, gas or water form?*).
2. If the reaction occurs, balance the equation.
3. On the watch glass, combine a few drops of each reactant. Record the observations for each reaction.

1) sodium phosphate + magnesium nitrate **Observations:**

2) copper (II) sulfate + barium nitrate **Observations:**

3) potassium chloride + silver (I) nitrate **Observations:**

4) sodium chloride + calcium nitrate **Observations:**

5) sodium phosphate + silver (I) nitrate **Observations:**

6) potassium chloride + sodium hydroxide **Observations:**

7) copper (II) sulfate + calcium nitrate **Observations:**