

# "Chapter 11 - Balancing Chemical Equations"

Key

## Chapter 7: Chemical Equations

Name: \_\_\_\_\_

Date: \_\_\_\_\_

Write and balance the following chemical equations.

1. Bromine gas reacts with aqueous potassium iodide to form potassium bromide and iodine gas.



2. Solid cobalt reacts with oxygen gas to produce a solid cobalt (III) oxide.



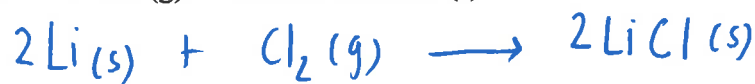
3. Phosphorous pentabromide reacts with water to yield phosphoric acid and hydrobromic acid.



4. Copper (II) oxide + sulfuric acid → copper (II) sulfate + water



5. Lithium (s) + chlorine (g) → lithium chloride (s)



6. Nickel (II) sulfide (s) + oxygen (g) → nickel (II) oxide (s) + sulfur dioxide (g)



7. Calcium hydride (s) + water (l) → calcium hydroxide (s) + hydrogen (g)



8. Diboron trioxide (s) + carbon (s) → tetraboron tricarbide (s) + carbon dioxide (g)



9. Silver (I) nitrate (aq) + zinc (s) → zinc (II) nitrate (aq) + silver (s)



10. Sodium chloride (s) + sulfuric acid (aq) → hydrochloric acid (aq) + sodium sulfate (s)

