

Write and balance the following chemical equations.

1. Bromine gas reacts with aqueous potassium iodide to form potassium bromide and iodine gas.
2. Solid cobalt reacts with oxygen gas to produce a solid cobalt (III) oxide.
3. Phosphorous pentabromide reacts with water to yield phosphoric acid and hydrobromic acid.
4. Copper (II) oxide + sulfuric acid \rightarrow copper (II) sulfate + water
5. Lithium (s) + chlorine (g) \rightarrow lithium chloride (s)
6. Nickel (II) sulfide (s) + oxygen (g) \rightarrow nickel (II) oxide (s) + sulfur dioxide (g)
7. Calcium hydride (s) + water (l) \rightarrow calcium hydroxide (s) + hydrogen (g)
8. Diboron trioxide (s) + carbon (s) \rightarrow tetraboron tricarbide (s) + carbon dioxide (g)
9. Silver (I) nitrate (aq) + zinc (s) \rightarrow zinc (II) nitrate (aq) + silver (s)
10. Sodium chloride (s) + sulfuric acid (aq) \rightarrow hydrochloric acid (aq) + sodium sulfate (s)