



UNIT 3 – BONDING

IPOD Questions



IT'S *THE* PROBLEM OF *THE* DAY

IPOD # 14

NAME OR WRITE A FORMULA FOR THE FOLLOWING EXAMPLES...

- NaOH
- Beryllium Sulfate
- Tin(II) Iodide
- Aluminum Cyanide
- Zinc Hydroxide
- Co_3N_2
- Mg_3P_2
- Beryllium Acetate
- Fe_3N_2
- $\text{Ga}(\text{NO}_2)_3$
- Silver Sulfide
- Ag_2SO_3



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IPOD # 15

Show how an ionic bond forms by completing the following:

Metal	Valence shell	e⁻ dot symbol	Ion formed	Valence shell of ion	Chemical name and formula of compound
aluminum					
Nonmetal	Valence shell	e⁻ dot symbol	Ion formed	Valence shell of ion	
bromine					

Bonding model

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IPOD # 16

NAME OR WRITE A FORMULA FOR THE FOLLOWING EXAMPLES...

- K_3PO_4
- Sodium nitride
- Boron trichloride
- HI
- N_2O_5
- Calcium sulfate
- Acetic acid
- $HBrO_2$
- Lead (II) hydroxide
- Sulfuric acid
- KBr
- CCl_4
- Cu_2SO_4
- Dinitrogen tetrahydride



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IPOD # 17

DRAW A LEWIS DOT STRUCTURE FOR THE FOLLOWING...

- CH_2Cl_2
- Dihydrogen monosulfide
- Nitrogen trichloride
- Lead (II) hydroxide
- Boron triiodide
- Carbon dioxide



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IPOD # 18

COMPLETE THE FOLLOWING TABLE...

<u>Name</u>	<u>Formula</u>	<u>Lewis Dot Structure</u>	<u>Molecular Shape</u> VSEPR Shape	<u>Bond Polarity</u> Use EN differences to calculate	<u>Molecule Type</u> polar or nonpolar based on molecule symmetry	<u>Intermolecular Attractions</u> London dispersion, dipole, hydrogen bonding?
silicon tetrachloride						
Dihydrogen monoselenide						
	NH ₃					