
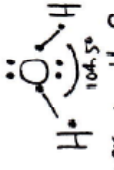

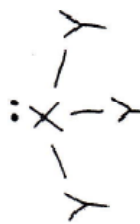
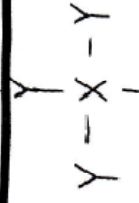

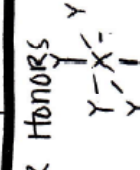


Atoms bonded to central atom	# of unshared pairs of e ⁻ around central atom	Shape (geometry)	ewis dot diagram (sample (s))	Bond angle	* NO. (s) • Elements want 8e ⁻ • H only wants 2e ⁻ • line, — 2 shared e ⁻ (single bond) • put element that there is only 1 of in the center. • line, = 4 shared e ⁻ (double bond) • unshared pairs take up more space than shared pairs • B can have 6e ⁻ * Exception *
X — Y * no central atom *	N/A	N/A	H — F: other ex → O ₂ , N ₂ , Cl ₂	180°	
X — Y — X * 2 A	none	Linear	:O=Si=O: 180°	180°	
	2 pairs	Bent	 other ex → H ₂ S	104.5°	
	∅	Trigonal planar	BI ₃ 120°	120°	
	1 pair	Trigonal pyramidal	NCl ₃ 105°	105°	
	∅	Tetrahedral	CB ₄ 109.5°	109.5°	• Carbon is always in the center
	∅	Trigonal Bipyramidal	PCl ₅	120° 90°	• S or P can have 10e ⁻ * Exception *
FOR HONORS 	∅	Octahedral	SF ₆	90°	