

Name _____ Date _____

PERIODIC TRENDS #3

| I | II | III | IV | V | VI | VII | VIII |
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Use the code letters A to Z and the clues below to fill in the periodic table above.

- The following elements belong in the same families
BFT---DGLZ JNV CMS QXY AEO IPH UKWR
- The charge on an ion of H is +4 and -4
- A neutral atom of C contains 8 electrons
- G is a noble gas
- U is an alkali metal
- The electrons of E end with s^2p^3 (E has 5 electrons in its outermost shell)
- N is an alkaline earth metal
- T has 3 electrons in the fourth energy level
- Q is a halogen
- F has the smallest atomic mass in its family
- T is more metallic than B
- J has 20 protons
- P has the lowest ionization energy in its family
- N has 5 neutrons in its nucleus
- S's atomic radius is greater than that of C
- C is more closely related to S than to M
- Y is a liquid whereas Q is a gas
- X has the highest electronegativity in its family
- K has the largest radius in its family
- W is a gas
- R has a higher ionization energy than U
- Atom Z has 2 neutrons
- D contains 10 protons
- The electrons of atom G are distributed over 3 energy levels
- H is the least metallic element in its group
- The electrons of atom O are distributed over 3 energy levels
- A is more metallic than either O or E

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