

Good Morning!!!

- I am handing the homework back. Please take a look at the answer keys on the side of the room.
- Place a hash mark (||) not a hash tag (#) by any problems that you would like to be discussed during the review.

Lab Review

- Come take a look at the beaker where you all put your material from yesterday's lab.
- Think about your observations in terms of electron exchanges.
- With your partner, at your desk draw a picture of this exchange. There are supplies up front. Be colorful. They may end up on the wall.

Share

- Get up and go to a lab group that is not at your bench.
- Present your picture to the other group.
- What differences might you have?

Considerations

- Were there any materials that did not fully react? Why is this so?
- How can I isolate the precipitate?
- Where did the aluminum go?
- Can I isolate it?

Percent Mass

- If I assume that everyone (all 13 groups) used 2 grams of copper chloride, how much copper metal would should have formed.

Most Requested Homework Topics

Core I: Periodicity

- Take a look at your core.
- Make sure that the correct grade is on IC.
- I will take questions in a moment.

Electrolysis

- Electricity is passed through an aqueous solution.
- There is an electron exchange at the anode and cathode.
- Recall the light bulb.

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- Click physical science
- Grades 9-12: Chemical reactions
- #3: Electrolysis: Use your understanding of electron exchange to predict the greatest mass exchange.

Answer the following on your whiteboards.

- As electrons move from the black wire into the solution, what happens?
- What did the color changes on the anodes and cathodes mean?

Batteries

- Anode/Cathode
- Electrolyte (salt bridge)

