

Worksheet #4: Single-Replacement Reactions

Step 1 - Write the formulas of the reactants on the left of the yield sign

Step 2 - Look at the Activity Series on page 333 to determine if the replacement can happen

Step 3 - If the replacement can occur, complete the reaction and balance it. If the reaction cannot happen, write N.R. (no rxn) on the product side.

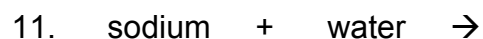
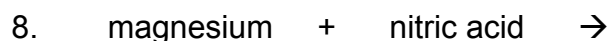
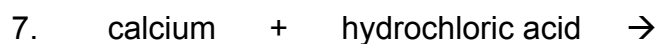
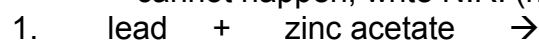


Table 11.2

Activity Series of Metals

| | Name | Symbol |
|----------------------------|------------|--------|
| Decreasing reactivity ↓ | Lithium | Li |
| | Potassium | K |
| | Calcium | Ca |
| | Sodium | Na |
| | Magnesium | Mg |
| | Aluminum | Al |
| | Zinc | Zn |
| | Iron | Fe |
| | Lead | Pb |
| | (Hydrogen) | (H)* |
| | Copper | Cu |
| | Mercury | Hg |
| Silver | Ag | |

*Metals from Li to Na will replace H from acids and water; from Mg to Pb they will replace H from acids only.

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