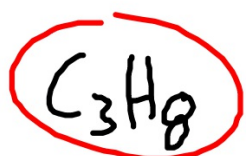


% Composition (mass)

H \rightarrow 18% C \rightarrow 82%



% C 12g/mol 12g(3) = 36.0g C $\left. \begin{array}{l} \\ \end{array} \right\} 44$

% H 1.0g/mol 1g(8) = 8g H

$\frac{36.0g}{44.0g} = 82\%$ $\frac{8g}{44g} = 18\%$



$$\text{N} = 14 \times 2 = 28$$

$$\text{H} = 1 \times 4 = 4$$

$$\text{O} = 16 \times 3 = 48$$

$$80$$

$$\frac{28}{80} \times 100 = 35\%$$

$$\frac{4}{80} \times 100 = 5\%$$

$$\frac{48}{80} \times 100 = 60\%$$

