

3.12.20

PT: Periodic Trends, pt2

Today's Objectives:

- **Recap Group and Period**
- **Learn electronegativity**
- **Learn reactivity**

Mendeleev: *invents periodic table so people won't have to memorize the elements' properties*

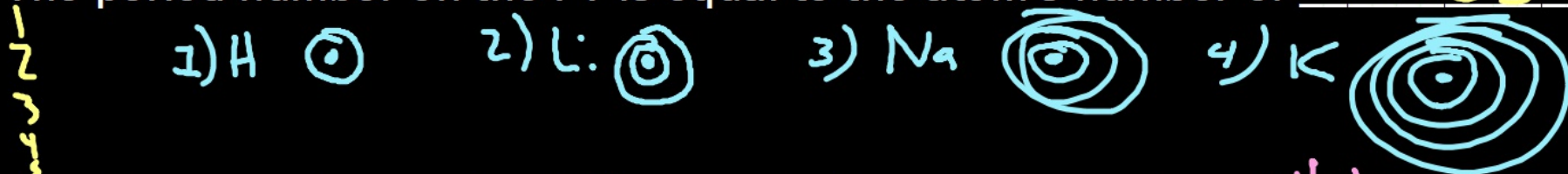
Chemistry teachers: *makes the students memorize the table*

Mendeleev:



- Which element is the largest? Francium (Fr)
- Which element is the smallest? Fluorine (F)

The period number on the PT is equal to the atom's number of Energy Levels



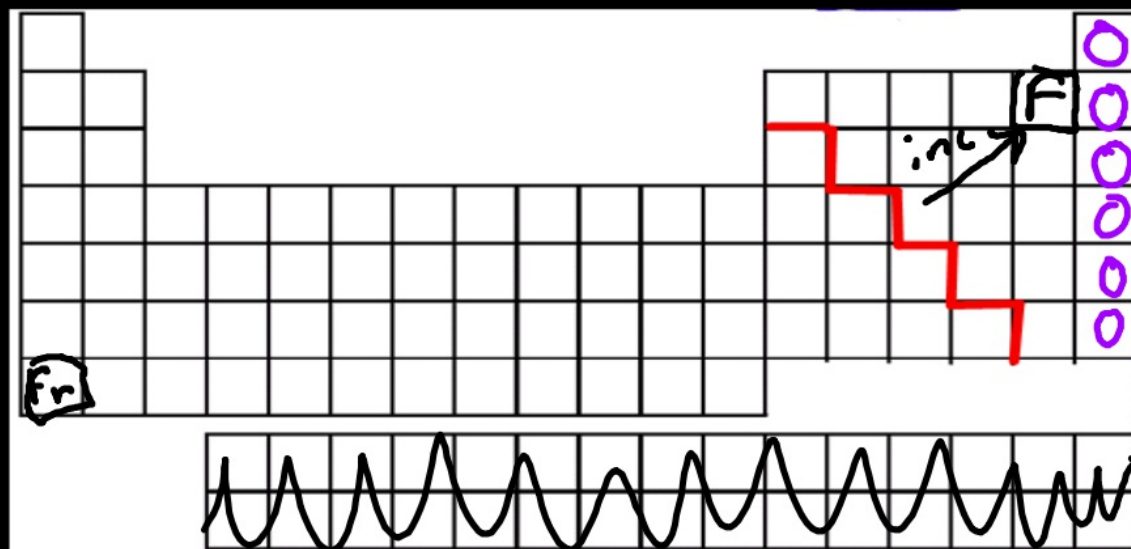
The group number on the PT is equal to the atom's number of Valence Electrons.

← 1 2 3 4 5 6 →

What ions do these elements form?

C	N	Br	S	Al	Kr	Fr
-4	-3	-1	-2	+3	0	+1

Electronegativity - How badly the atom wants
more valence electrons.



Least =
Francium →

Most
Electronegative = Fluorine!

Noble Gases -
Already
have a
completed
octet!
(8)

Reactivity - how easily an element enters a chemical reaction.

What is better for jewelry?
A reactive or non-reactive metal?

Doesn't rust!
:)

Au, Ag

Ion Formation Vocab

Metals Lose electrons to form + ions called Cations.

Non-metals Gain electrons to form - ions called Anions.