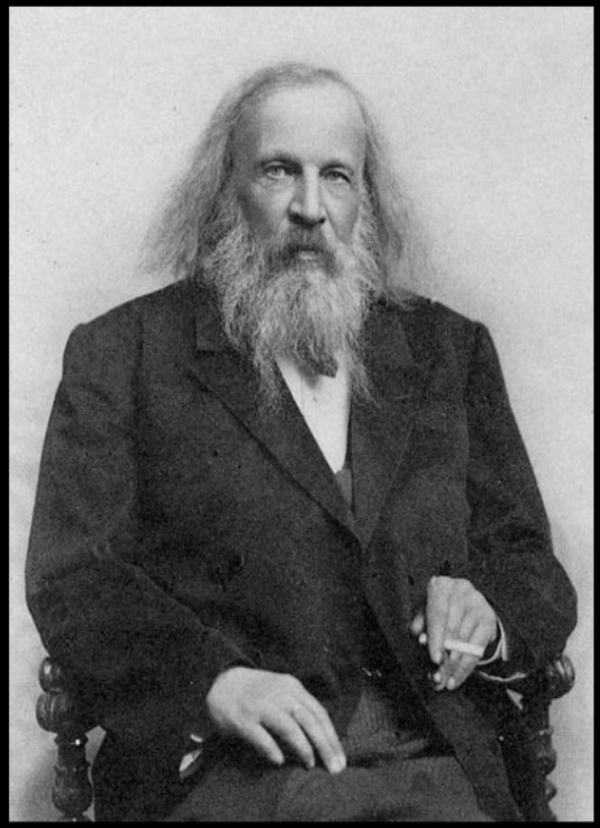
3.10.20

Periodic Table: Intro

Today's Obejctives:

- Learn the history of the PT
- Know the Group Names
- Research Your Element!



Why is the PT arranged the way it is? Properties of elements.

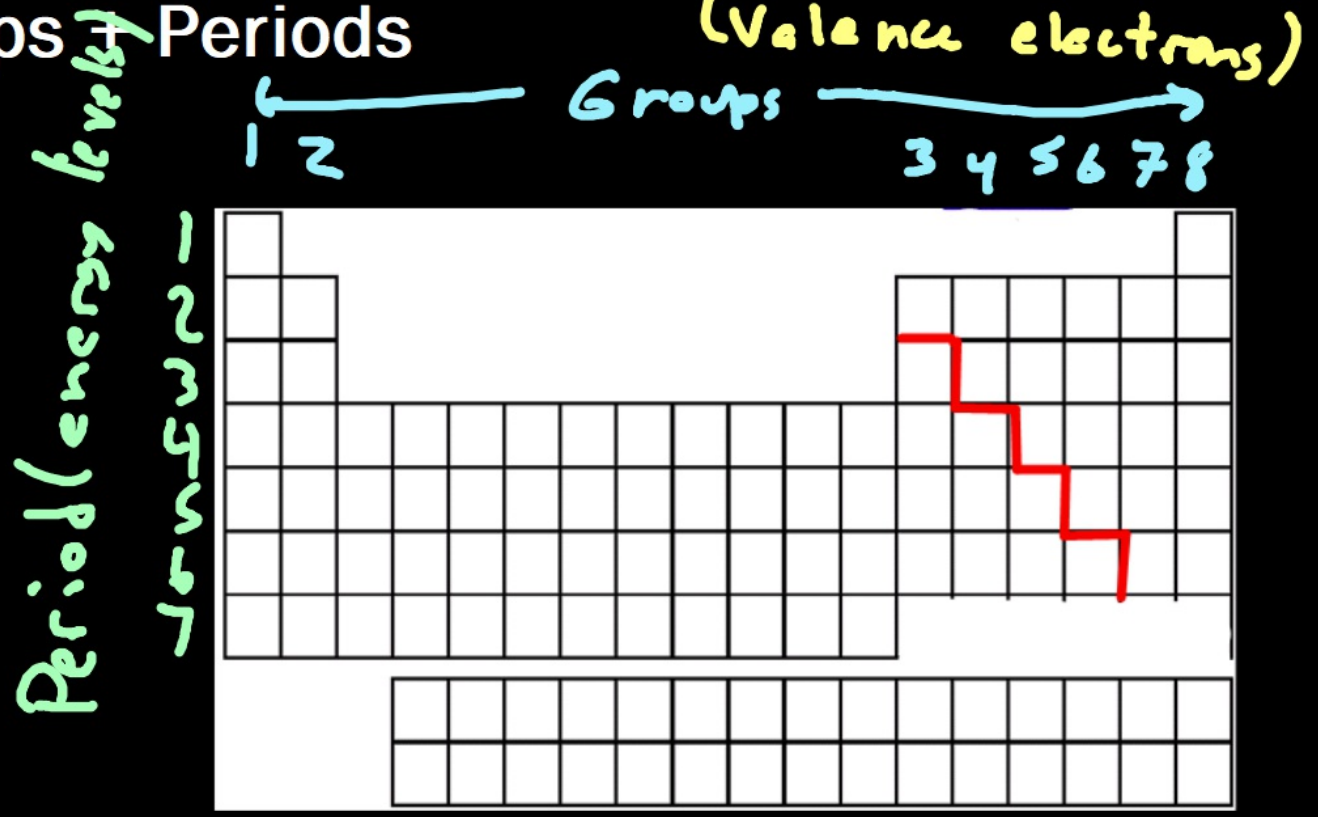
periodic law = Repeating set of element behavior.

① Mendeleev PT → ② Moseley PT:

increasing Mass

increasing atomic Number.

Groups → Periods



Which elements behave most like (fluorine)?

Cl, Br, I, At

Group Names to Know:

1: Alkali Metals (Li → Fr)

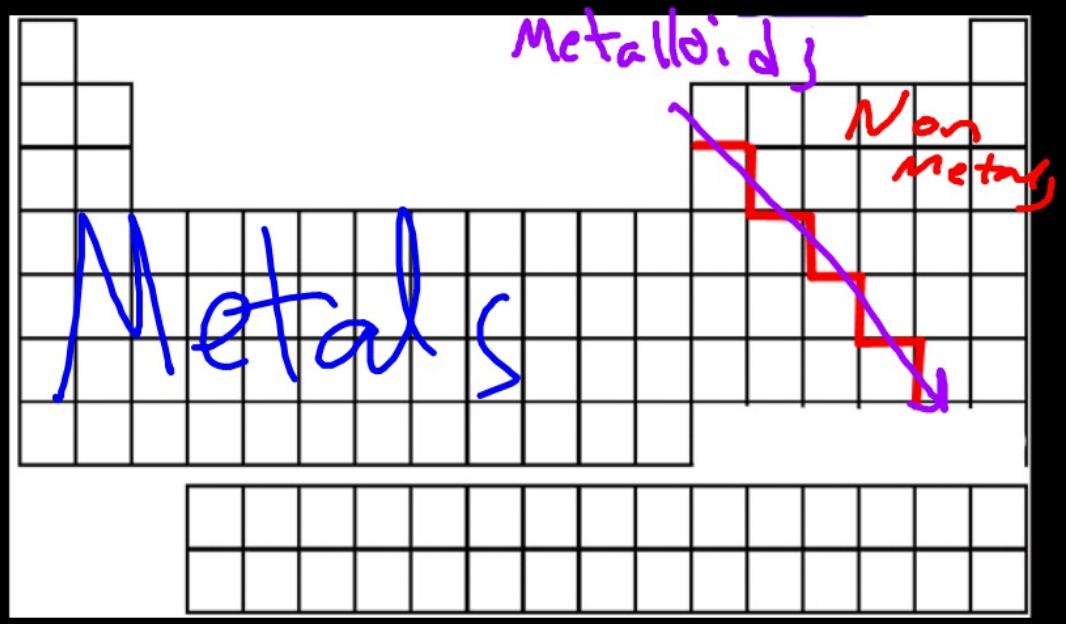
2: Alkaline Earth Metals (Be → Ra)

Middle of PT: Transition Metals

7: Halogens (F → At)

8: Noble Gases (He → Rn)

Metals/Metalloids/Non-Metals



Metals

- Solids at Room Temp
- High Melting Points
- Good conductors of light & heat
- Malleable (foil)
- Ductile (wire)

Non-Metals

- Liquids + Gases at R.T.
- Low

Not