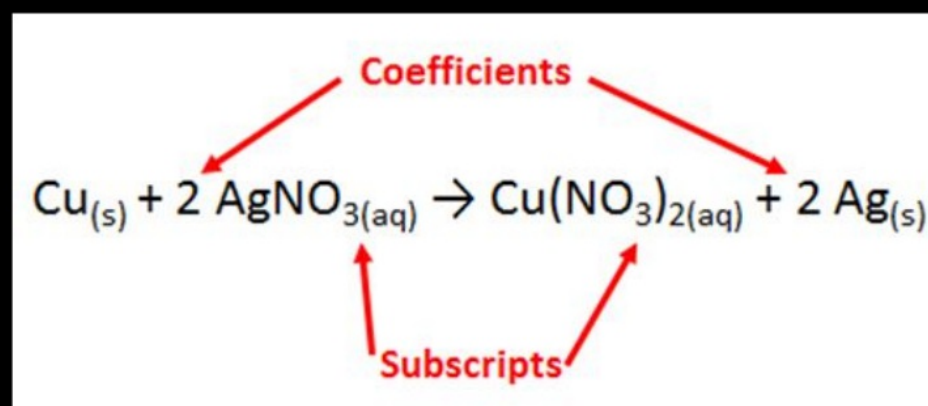


## 5.4.20

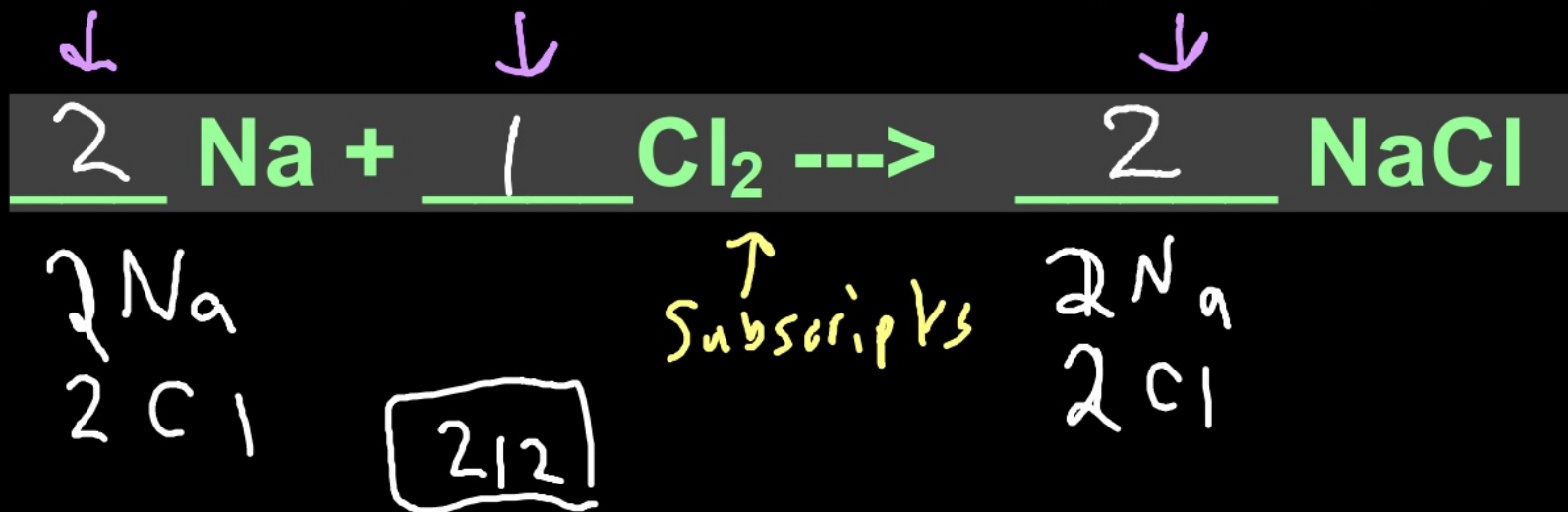
# Balancing Chemical Reactions

### Today's Objectives:

- What are coefficients?
- Work through balancing Qs



- All chemical reactions need to be balanced. (atoms on left of arrow = atoms on right)
- Balance them by adding a number (called coefficients) in front of each chemical.
- Do not change the little numbers (subscripts!)

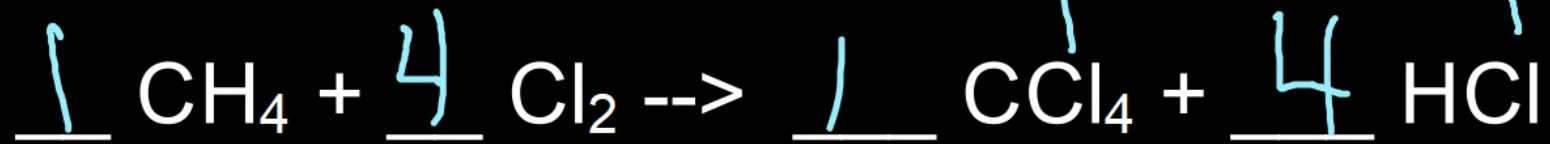


## Balancing Hints:

- Balance the metals first
- Put a 2 in front of an odd subscript
- If an element is by itself, save it for last!

(Ex: 3, 5)





1 C  
4 H  
2 Cl

1 C  
4 H  
8 Cl

1414



4P

20

4P

60

132